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- Electronic and atomic structure of matter
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- Solution chemistry
- Acids and bases
- Stoichiometry
- Equilibrium and reaction rates
- Thermochemistry
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Editorial Review

From the Author

Scoring well on College-Level Examination Program (CLEP) Chemistry is important for your ability to bypass taking the class and earn college credit. To achieve a high score, you need to develop skills to properly apply the knowledge you have and quickly choose the correct answer. You must solve numerous practice questions that represent the style and content of the CLEP questions. Understanding key science concepts is more valuable than memorizing terms.

This book provides **over 1,200 chemistry practice questions** that test your knowledge of all CLEP Chemistry topics. At the back of the book, you will find answer keys (use them to verify your answers) and explanations to questions. These explanations provide step-by-step solutions for quantitative questions and detailed explanations for conceptual questions. They also cover the foundations of general chemistry topics needed to answer related questions on the test. **By reading these explanations carefully and understanding how they apply to solving the question, you will learn important chemistry concepts and the relationships between them. This will prepare you for the test and you will significantly increase your score.**

All the questions are prepared by our science editors who possess extensive credentials and are educated in top colleges and universities. Our editors are experts on teaching sciences, preparing students for standardized science tests and have coached thousands of undergraduate and graduate school applicants on admission strategies.

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From the Inside Flap

Electronic and Atomic Structure of Matter; Periodic Table

- Atomic theory and atomic mass
- Atomic number and mass number; isotopes and mass spectroscopy
- Electron energy levels: atomic spectra, quantum numbers, atomic orbitals
- Periodic relationships (atomic radii, ionization energies, electron affinities, oxidation states)

Bonding

- Bonding properties and structures
- Ionic, covalent, metallic bonds
- Polarity of bonds, electronegativities
- Molecular geometry
- Intermolecular forces (hydrogen bonding, dipole-dipole, London dispersion forces)
- Molecular models: valence bond theory, hybridization of orbitals, resonance, *sigma* and *pi* bonds

States of Matter: Gases, Liquids, Solids

- Laws of ideal gases; equations of state for an ideal gas

- Kinetic molecular theory
- Molar volumes, density
- Types of solids, phase changes, phase diagrams
- Changes of state, critical phenomena, crystal structure
- Intermolecular forces

Solution Chemistry

- Solution preparation, types of solutions, stoichiometry
- Molarity and percent by mass concentrations
- Solubility of solids, liquids and gases
- Precipitation and solubility rules
- Colligative properties; interionic attraction

Acids and Bases

- Strong and weak acids and bases
- Brønsted-Lowry theory
- Concepts of Arrhenius and Lewis
- pH, titrations, indicators
- Oxidation-reduction reactions

Stoichiometry

- Molar mass, Avogadro's number
- Empirical and molecular formulas
- Ionic and molecular species present in chemical systems; net-ionic equations
- Balancing of chemical equations
- Stoichiometric calculations, percent yield, limiting reactants

Equilibrium and Reaction Rates

- Equilibrium systems
- Le Châtelier's principle and position of equilibrium
- Equilibrium constants, equilibrium expressions
- Factors affecting reaction rates; rate constant; rate-determining step
- Potential energy, activation energy

Thermodynamics

- First law: heat of formation; heat of reaction; change in enthalpy, Hess's law; heat capacity; heats of vaporization and fusion
- Second law: free energy of formation; free energy of reaction; dependence of change in free energy on enthalpy and entropy changes
- Chemical reactions, heating and cooling curves, entropy

Electrochemistry

- Galvanic and electrolytic cells
- Redox reactions (including half-reactions) within electrochemical cells
- Standard reduction potentials, oxidation number

Answer Keys & Explanations

From the Back Cover

Scoring well on College-Level Examination Program (CLEP) Chemistry is important for your ability to bypass taking the class and earn college credit. To achieve a high score, you need to develop skills to properly apply the knowledge you have and quickly choose the correct answer. You must solve numerous practice questions that represent the style and content of the CLEP questions. Understanding key science concepts is more valuable than memorizing terms.

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Users Review

From reader reviews:

Phillip Chadwick:

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